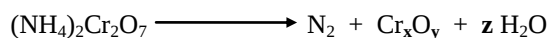


BALANCEAMENTO DE EQUAÇÕES

- A) $\text{C}_2\text{H}_6\text{O} + \text{O}_2 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$
- B) $\text{Na}_2\text{CO}_3 + \text{HCl} \longrightarrow \text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$
- C) $\text{C}_6\text{H}_{12}\text{O}_6 \longrightarrow \text{C}_2\text{H}_6\text{O} + \text{CO}_2$
- D) $\text{C}_4\text{H}_{10} + \text{O}_2 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$
- E) $\text{FeCl}_3 + \text{Na}_2\text{CO}_3 \longrightarrow \text{Fe}_2(\text{CO}_3)_3 + \text{NaCl}$
- F) $\text{NH}_4\text{Cl} + \text{Ba}(\text{OH})_2 \longrightarrow \text{BaCl}_2 + \text{NH}_3 + \text{H}_2\text{O}$
- G) $\text{Ca}(\text{OH})_2 + \text{H}_3\text{PO}_4 \longrightarrow \text{Ca}_3(\text{PO}_4)_2 + \text{H}_2\text{O}$
- H) $\text{Fe}_2(\text{CO}_3)_3 + \text{H}_2\text{SO}_4 \longrightarrow \text{Fe}_2(\text{SO}_4)_3 + \text{H}_2\text{O} + \text{CO}_2$
- I) $\text{Na}_2\text{O} + (\text{NH}_4)_2\text{SO}_4 \longrightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{NH}_3$
- J) $\text{FeS}_2 + \text{O}_2 \longrightarrow \text{Fe}_2\text{O}_3 + \text{SO}_2$
- K) $\text{NH}_3 + \text{O}_2 \longrightarrow \text{NO} + \text{H}_2\text{O}$
- L) $\text{KMnO}_4 + \text{H}_2\text{SO}_4 \longrightarrow \text{Mn}_2\text{O}_7 + \text{K}_2\text{SO}_4 + \text{H}_2\text{O}$
- M) $\text{CS}_2 + \text{O}_2 \longrightarrow \text{CO}_2 + \text{SO}_2$
- N) $\text{H}_3\text{PO}_4 + \text{CaO} \longrightarrow \text{Ca}_3(\text{PO}_4)_2 + \text{H}_2\text{O}$
- O) $\text{Na}_2\text{CO}_3 + \text{H}_3\text{PO}_4 \longrightarrow \text{Na}_3\text{PO}_4 + \text{H}_2\text{O} + \text{CO}_2$
- P) $\text{KMnO}_4 \longrightarrow \text{K}_2\text{MnO}_4 + \text{MnO}_2 + \text{O}_2$
- Q) $\text{Na} + \text{KNO}_3 \longrightarrow \text{Na}_2\text{O} + \text{K}_2\text{O} + \text{N}_2$
- R) $\text{Ni}(\text{CO})_4 \longrightarrow \text{Ni} + \text{CO}$
- S) $\text{CaC}_2 + \text{H}_2\text{O} \longrightarrow \text{C}_2\text{H}_2 + \text{CaO}$

QUESTÕES DE VESTIBULAR

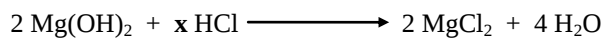
01 (FUVEST) A decomposição térmica de 1 mol de *dicromato de amônio* é representada pela equação:



Os valores de **x**, **y** e **z** são, respectivamente:

- a) 2, 3 e 4
- b) 2, 7 e 4
- c) 2, 7 e 8
- d) 3, 2 e 4
- e) 3, 2 e 8

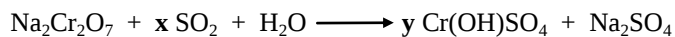
02 (ESAL/MG) A equação química:



fica estequiometricamente correta se **x** for igual a:

- a) 1
- b) 2
- c) 3
- d) 4
- e) 5

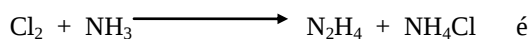
03 (PUCCAMP) Num “sapato de cromo”, o couro é tratado com um banho de “licor de cromo”, preparado através da reação representada pela equação:



Depois de balanceada com os menores coeficientes inteiros possíveis, ela apresenta:

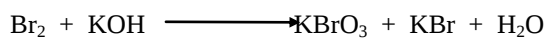
- | | x | y |
|----|----------|----------|
| a) | 3 | 2 |
| b) | 2 | 3 |
| c) | 2 | 2 |
| d) | 3 | 3 |
| e) | 2 | 1 |

04 (UNIP/SP) A soma dos menores coeficientes inteiros que balanceiam a equação:



- a) 4
- b) 15
- c) 21
- d) 8
- e) 6

05 (OSEC/SP) A soma dos coeficientes da equação abaixo é igual a



- a) 13
- b) 20
- c) 19
- d) 15
- e) 18

GABARITOS

- | | | | |
|----|---------------|----|----------------|
| A) | 1, 3, 2, 3 | K) | 4, 5, 4, 6 |
| B) | 1, 2, 2, 1, 1 | L) | 2, 1, 1, 1, 1 |
| C) | 1, 2, 2 | M) | 1, 3, 1, 2 |
| D) | 2, 13, 8, 10 | N) | 2, 3, 1, 3 |
| E) | 2, 3, 1, 6 | O) | 3, 2, 2, 3, 3 |
| F) | 2, 1, 1, 2, 2 | P) | 2, 1, 1, 1 |
| G) | 3, 2, 1, 6 | Q) | 10, 2, 5, 1, 1 |
| H) | 1, 3, 1, 3, 3 | R) | 1, 1, 4 |
| I) | 1, 1, 1, 1, 2 | S) | 1, 1, 1, 1 |
| J) | 4, 11, 2, 8 | | |

01 A 02 D 03 A 04 D 05 E